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unchanged. Rank the following pressures in increasing order. (c) 76 torr (a) 50 kPa O, OOÌctbv-xHome - Kenilworth Public SchoolsIf a gas and a liquid are the same temperature and pressure, diffusion occurs much faster in the gas because. A. there are more elastic collisions between the particles in a gas. B. gases are more compressible. C. the particles move faster in a gas and there is a greater distance between them.Chapter 11 Gases Review Flashcards | QuizletSection 11.4 Dalton's Law of Partial Pressures Goals To describe the properties of mixtures of gases. To describe calculations that deal with mixtures of gases. In the real world, gases are usually mixtures. This section describes how mixing gases affects the properties of the resulting mixture.Chapter 11 - Gases462 Chapter 11 Gases Discovering the Relationships Between Properties If we want to explain why a weather balloon carrying instruments into the upper atmosphere expands as it rises, we need to consider changes in the properties of the gases (pressure, volume, temperature, or number of gas particles) inside and outside the balloon.Chapter 11 Gases - An Introduction to ChemistryCHAPTER 11 REVIEW Gases SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. b Pressure surf f a o c r e ce area. For a constant force, when the surface area is tripled the pressure is (a) doubled. (b) a third as much. (c) tripled. (d) unchanged. 2. d, c, a, b Rank the following pressures in increasing order. (a) 50 kPa (c) 76 torr (b) 2 atm (d) 100 N/m2 3.mc06se cFMsr i-vi - Ed W. Clark High SchoolStart studying Chapter 11- Gases: Section 1: Gases and Pressure. Learn vocabulary, terms, and more with flashcards, games, and other study tools.Chapter 11- Gases: Section 1: Gases and Pressure ...this theory explains some of the properties of ideal gases. In this chapter, you will study the predictions of kinetic-molecular theory for gases in more detail. This includes the relationship among the temperature, pressure, volume, and amount of gas in a sample. SECTION 11.1 Key Te r m s pressure newton barometer millimeters of mercurySECTION 11.1 Gases and Pressure - Pickford High SchoolCHAPTER 11 REVIEW Gases SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. State whether the pressure of a fixed mass of gas will increase, decrease, or stay the same in the following circumstances: increase a. temperature increases, volume stays the same decrease b. volume increases, temperature stays the samemc06se cFMsr i-vi - Ed W. Clark High SchoolDownload chapter 11 review gases section 2 answers - Bing book pdf free download link or read online here in PDF. Read online chapter 11 review gases section 2 answers - Bing book pdf free download link book now. All books are in clear copy here, and all files are secure so don't worry about it.Chapter 11 Review Gases Section 2 Answers - Bing | pdf ...Chapter 11 Review Gases Section 2 Answers.pdf - search pdf books free download Free eBook and manual for Business, Education,Finance, Inspirational, Novel, Religion, Social, Sports, Science, Technology, Holiday, Medical,Daily new PDF ebooks documents ready for download, All PDF documents are Free,The biggest database for Free books and documents search with fast results better than any online ...Chapter 11 Review Gases Section 2 Answers.pdf | pdf Book ...Modern Chemistry 93 Gases CHAPTER 11 REVIEW Gases SECTION 1 SHORT

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If a gas and a liquid are the same temperature and pressure, diffusion occurs much faster in the gas because. A. there are more elastic collisions between the particles in a gas. B. gases are more compressible. C. the particles move faster in a gas and there is a greater distance between them.

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