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 1. Heat - energy transfer resulting from a potential energy difference between system and surroundings. 2. Work - energy transfer resulting from a force exerted through a distance by the system to the surroundings. Energy storage and energy transfer is limited to these and only these forms in classical thermodynamics.
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 Chapter 3: Energy Transport by Heat, Work, and Mass
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 The transfer of energy between a chemical reaction system and its surroundings occurs and work or heat. ΔU (or ΔE) is the change in internal energy of the system q is heat and w is work
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 Heat (Energy) The SI-unit of heat - or energy - is joule (J).. With temperature difference . heat will transfer from a warm body with higher temperature to a colder body with lower temperature; Other units used to quantify heat are the British Thermal Unit - Btu (the amount of heat to raise 1 lb of water by 1 °F) and the Calorie (the amount of heat to raise 1 gram of water by 1 °C (or 1 K)).
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 2Ch 3 energy transfer by work, heat and mass
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