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ANSWER Answer the
following questions in
the space provided. 1.
b The coefficients in a
chemical equation
represent the (a)
masses in grams of all
reactants and
products. (b) relative
number of moles of
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9-3 PROBLEMS Write
the answer on the line
to the left. Show all
your work in the space
provided. 1. 88% If the
actual yield of a
reaction is 22 g and
the theoretical yield is
25 g, calculate the
percent yield. 2. 6.0
mol of N₂ are mixed
with 12.0 mol of H₂
according to the
following equation: N₂(g)
+ 3H₂(g) → 2NH₃(g)
2.0 mol
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. SECTION 2.
PROBLEMS Write the
answer on the line to
the left. Show all your
work in the space
provided. 1. The

following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(s) \rightarrow 2\text{KCl}(s) + 3\text{O}_2(g)$...

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questions in the space provided. 1. ____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products.

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 Composition
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 mass relationships of
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 Reaction stoichiometry
 -The mass
 relationships between
 reactants and products

in a chemical reaction
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pter 9 focuses on
reaction stoichiometry:
using a balanced
chemical equation to
calculate the number
of grams, moles, or
particles of
reactants/products
involved in a chemical
reaction. Students had
an introduction to
composition
stoichiometry in
Chapter 3 and will now
move on to some more
difficult
problems.Chapter 9 -
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 following equation
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 preparation for oxygen
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above we can see that
 if we have 12.4 mol H
 2 we need 4.13 mol N
 2. We don't have that
 much N 2 so the .892
 mol of N 2 must be the
 limiting reagent. We
 can now determine
 how much ammonia
 will be produced using
 the mole ratio in the
 balanced equation
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 9 - Stoichiometry
 Review #1 - #18, #31,
 & #38 Answers . 38. To
 ensure that all
 magnesium is
 converted to MgO, I
 would use pure
 oxygen, not air, to
 carry out the reaction,
 because Mg could
 react with N₂ in air to
 form Mg₃N₂. The pure
 oxygen should be in
 excess. 5. a. 50 mol HI
 6. a. 15.8
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1 SHORT ANSWER
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SECTION 2. PROBLEMS
Write the answer on
the line to the left.
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Chapter 9 describes
how to use mole ratios,
molar masses,
conversions, limiting
reactants, and percent
yield to ...
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Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

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Stoichiometry SHORT ANSWER Answer the following questions in the space provided. 1. b The coefficients in a chemical equation represent the (a masses in grams of all reactants and products.

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mol of N 2 are mixed

with 12.0 mol of H 2
according to the
following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$
N 2; 2.0 mol a.
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From above we can
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12.4 mol H 2 we need
4.13 mol N 2. We don't
have that much N 2 so
the .892 mol of N 2
must be the limiting
reagent. We can now
determine how much
ammonia will be
produced using the
mole ratio in the
balanced equation :
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-The mass

relationships between

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in a chemical reaction

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