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1.8. Why are ionic solids conducting in the molten state and not in

the solid state ? Ans: In the ionic solids, the electrical conductivity is due to the movement of the ions. Since the ionic mobility is negligible in the solid state, these are non-conducting in this state.
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 lattice and atoms of X
 occupy 1/3rd of
 tetrahedral voids? Ans. 2.
 What is the formula of a
 compound [...]Important
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as in molten state and
 melts at extremely high
 temperature. What type
 of solid is it? Answer:
 Given solid 'A' is a
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element Y forms ccp

lattice and atoms of X

occupy 1/3rd of

tetrahedral voids? Ans. 2.

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Question: 1.7 - Solid A is a very hard electrical insulator in solid as well as in molten state and melts at extremely high temperature. What type of solid is it? Answer: Given solid 'A' is a covalent solids, such as diamond. Question: 1.8 - Ionic solids conduct electricity in molten state but not in solid state. Explain.

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