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 Acids and Bases . 14.1 The Nature of Acids
 and Bases . A. Arrhenius Model 1. Acids
 produce hydrogen ions in aqueous
 solutions 2. Bases produce hydroxide ions
 in aqueous solutions B. Bronsted-Lowry
 Model 1. Acids are proton donors 2. Bases

are proton acceptors 3. H₃O⁺ is called the hydronium ion C. Conjugate Acid-Base Pairs 1. Chapter 14 - Acids and Bases - ScienceGeek.net CHAPTER 14 REVIEW Acids and Bases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Answer the following questions according to the Brønsted-Lowry definitions of acids and bases: HSO₃⁻ a. What is the conjugate base of H₂SO₃? NH₃ b. What is the conjugate base of NH₄⁺?

14 Acids and Bases Acids and Bases Know the definition of Arrhenius, Bronsted-Lowry, and Lewis acid and base. Autoionization of Water Since we will be dealing with aqueous acid and base solution, first we must examine the behavior of water. Chapter 14: Acids and Bases Acids and Bases: Chapter 14 & 15 . HW: •Read Ch 14: •Fill in as much of the acid base table as you can, as you read . Acid base conductivity and reactivity Conductivity, Reactivity,, Hydrochloric*acid* high high Acidic*acid* ** low* medium Disalld*water* none* none* Ammoniumhydroxide med* none* Sodiumhydroxide* high none* *Acids and Bases: Chapter 14 & 15 Learn chemistry chapter 14 acids bases with free

interactive flashcards. Choose from 500 different sets of chemistry chapter 14 acids bases flashcards on Quizlet. chemistry chapter 14 acids bases Flashcards - Quizlet CHAPTER 14 REVIEW . Acids and Bases. SHORT ANSWER Answer the following questions in the space provided. 1. a. Write the two equations that show the two-stage ionization of sulfurous acid in water. stage 1: H₂SO₃(aq) + H₂O(l) ~ H₃O⁺(aq) + HSO₃⁻(aq) b. CHAPTER 14 REVIEW Acids and Bases - Weebly Learn acids and bases chapter 14 chemistry with free interactive flashcards. Choose from 500 different sets of acids and bases chapter 14 chemistry flashcards on Quizlet. acids and bases chapter 14 chemistry Flashcards - Quizlet Arrhenius Acid-Base Theory Arrhenius Acid -a hydrogen containing compound that ionizes to produce hydrogen ions (H⁺) when dissolved in water Because molecular acids are not made of ions, they cannot dissociate. HCl (aq) → H⁺(aq) + Cl⁻(aq) They must be pulled apart, or ionized, by the water. Chapter 14: Acids and Bases Ch 14 Page 1 Chapter 14: Acids and Bases CHAPTER 14 ACIDS AND BASES 5 When H₃PO₄ is added

to water, the three acids that are present are H₃PO₄, H₂PO₄⁻, and HPO₄²⁻. H₃PO₄, with the largest K_a value, is the strongest of these weak acids. The conjugate bases of the three acids are H₂PO₄⁻, HPO₄²⁻

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Bases.ppt — application/vnd.ms-powerpoint, 5.30 MB (5561856 bytes)Chapter 14 - Acids and Bases — HCC Learning WebChapter 14. Acid-Base Equilibria. 14.1 Brønsted-Lowry Acids and Bases Learning Objectives. By the end of this section, you will be able to: Identify acids, bases, and conjugate acid-base pairs according to the Brønsted-Lowry definition; Write equations for acid and base ionization reactions;14.1 Brønsted-Lowry Acids and Bases - ChemistryA vodcast I made for watching later in the chapter (after we have talked about strong and weak acids and bases in class as well as salts). It goes over all the different categories of pH problems in this chapter (strong acid, weak base, salt of weak acid/strong base, etc) and how to solve them.Chapter 14 - Acids and Bases - Mrs. Duffey - FHNMajor topics: types of acids, amphoterism, pH scale, simple pH calculations, strong acid calculations, weak acid calculations (ICE tables), & acid mixture problems. Learn chemistry chapter 14 acids bases with free interactive flashcards. Choose from 500 different sets of chemistry chapter 14 acids bases flashcards on

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Major topics: types of acids, amphoterism, pH scale, simple pH calculations, strong acid calculations, weak acid calculations

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Chapter 14: Acids and Bases

CHAPTER 14 REVIEW . Acids and Bases.

SHORT ANSWER Answer the following questions in the space provided. 1. a.

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Arrhenius Acid-Base Theory Arrhenius Acid -a hydrogen containing compound that ionizes to produce hydrogen ions (H⁺) when dissolved in water Because molecular acids are not made of ions, they cannot dissociate. HCl (aq) → H⁺(aq) + Cl⁻(aq) They must be pulled apart, or ionized, by the water. Chapter 14: Acids and Bases Ch 14 Page 1

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Chapter 14 - Acids and Bases - Mrs. Duffey - FHN

Acids and Bases: Chapter 14 & 15 . HW:

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Chapter 14 - Acids and Bases

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CHAPTER 14 REVIEW Acids and Bases

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