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# Chapter 16 Review Acid Base Titration And Ph 2

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always produced by a neutralization reaction? a. acid and a base b. water and a base c. water and an acid d. water and a salt

2. Chapter 16: Leader: Acid and Base Review Chapter 16 Acids and Bases

1. Acids were recognized primarily from their sour taste. Bases were recognized from their bitter taste and slippery feel on skin.

2. In the Arrhenius definition, an acid is a substance that produces hydrogen ions ( $H^+$ ) when dissolved in water, whereas a base is a substance that produces hydroxide ions ( $OH^-$ ) in water.

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Chapter 16 Acid-Base ... 16.1 Acids and Bases: A Brief Review

- Acids taste sour and cause certain dyes to change color.
- Bases taste bitter and feel soapy.
- Arrhenius concept of acids and bases:
  - An acid is a substance that, when dissolved in water, increases the concentration of  $H^+$  ions.
  - Example:  $HCl$  is an acid.
  - An Arrhenius base is a substance that, when dissolved in water, increases the concentration of  $OH^-$  ions.

AP Chemistry— CHAPTER 16 STUDY GUIDE Acid-Base Equilibrium Acids and Bases Acid and Base Strength In any acid-base reaction, the equilibrium will favor the reaction that moves the proton to the stronger base.

$$HCl(aq) + H_2O(l) \rightarrow H_3O^+(aq) + Cl^-(aq)$$

$H_2O$  is a much stronger base than  $Cl^-$ , so the equilibrium lies so far to the right  $K$  is not measured ( $K \gg 1$ ).

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16.10: Acid-Base Behavior and Chemical Structure

Inductive effects and charge delocalization significantly influence the acidity or basicity of a compound. The acid-base strength of a molecule depends strongly on its structure. The weaker the  $A-H$  or  $B-H$  bond, the more likely it is to dissociate to form an  $H^+$  ion.

16: Acid-Base Equilibria - Chemistry LibreTexts CHAPTER 16 - Acid-Base Equilibria Section 16.1 - Acids and Bases: A Brief Review (a) Define an acid and a base, according to the Arrhenius definition. acid = base = (b) Write the products of each chemical reaction below, which involves the dissociation of each reactant into aqueous ions.

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- Acids and bases were first recognized by the properties of their aqueous solutions.
- o For example, acids turn litmus red, whereas bases turn litmus blue.

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16.10: Acid-Base Behavior and Chemical Structure Inductive effects and charge delocalization significantly influence the acidity or basicity of a compound. The acid-base strength of a molecule depends strongly on its structure. The weaker the A-H or B-H<sup>+</sup> bond, the more likely it is to dissociate to form an (H<sup>+</sup>) ion.

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Acid-Base ...

Chapter 16: Acid and Base Review Supplemental Instruction Iowa State University Leader: Kelsey Course: Chemistry 178 Instructor: Verkade Date: 10/10/2011 ~PLEASE DO NOT WRITE ON THIS WORKSHEET~

1. What two substances are always produced by a neutralization reaction? a. acid and a base b. water and a base c. water and an acid d. water and a salt

*AP Chemistry— CHAPTER 16 STUDY GUIDE Acid-Base Equilibrium*

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Acids and Bases Acid and Base Strength In any acid-base reaction, the equilibrium will favor the reaction that moves the proton to the stronger base.  $\text{HCl (aq)} + \text{H}_2\text{O(l)} \rightarrow \text{H}_3\text{O}^+(\text{aq}) + \text{Cl}^-(\text{aq})$   $\text{H}_2\text{O}$  is a much stronger base than  $\text{Cl}^-$ , so the equilibrium lies so far to the right  $K >> 1$ .

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**Chapter 16 Acids and Bases**

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Arrhenius acids and bases: -- acid: an  $\text{H}^+$  donor  $\text{HA} \rightleftharpoons \text{H}^+(\text{aq}) + \text{A}^-(\text{aq})$  -- base: an  $\text{OH}^-$  donor  $\text{MOH} \rightleftharpoons \text{M}^+(\text{aq}) + \text{OH}^-(\text{aq})$  - Brønsted-Lowry acids and bases: -- acid: an  $\text{H}^+$  donor  $\text{HA} \rightleftharpoons \text{H}^+(\text{aq}) + \text{A}^-(\text{aq})$  -- base: an  $\text{H}^+$  acceptor  $\text{HB}$

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16.1 Acids and Bases: A Brief Review •Acids taste sour and cause certain dyes to change color. •Bases taste bitter and feel soapy. •Arrhenius concept of acids and bases: •An acid is a substance that, when dissolved in water, increases the concentration of  $\text{H}^+$  ions. •Example:  $\text{HCl}$  is an acid. •An Arrhenius base is a substance that, when dissolved in water, increases the concentration of  $\text{OH}^-$  ions.

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titration graph and

answer four questions  
about the data.

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