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*Determination of
normality of the
unknown Sodium
hydroxide solution 2.*

*Comparison of the
normality of the two
given sodium*

carbonate solutions 3.

Estimation of oxalic

acid Qualitative

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Nitrate 2. Copper

Sulphate 3. Zinc

Sulphate 4. XI

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whose strength is known is called a standard solution. Question.2.What is a normal solution ? Answer. A solution containing one gram-equivalent mass of the solute per litre of the solution is called [...]

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method or procedure for working out the titre or concentration of an analyte in a solution. This is done by measuring the volume of a standard solution of an appropriate reagent whose precise concentration is already known. Volumetric analysis, Daniele Naviglio « Analytical ...colorimetry, calorimetry, chromatography, calibration procedures and laboratory safety. Through the practical tasks in the unit, you will develop proficiency in the quantitative analytical techniques of titration and colorimetry, including learning to calculate the concentration of solutions. You will Unit 2: Practical Scientific Procedures and

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Answers Viva

Questions with answers
Question.1.What is a
standard solution ?

Answer. A solution
whose strength is
known is called a
standard solution.

Question.2.What is a
normal solution ?

Answer. A solution
containing one gram-
equivalent mass of the
solute per litre of the
solution is called [...]

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any method of
quantitative chemical
analysis in which the
amount of a substance
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measuring the volume that it occupies or, in broader usage, the volume of a second substance that combines with the first in known proportions, more correctly called titrimetric analysis (see titration).

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 (a) To prepare 250ml of M/20 solution of Mohr’s salt. (b) Using this calculate the molarity and strength of the given $KMnO_4$ solution. APPARATUS AND CHEMICALS REQUIRED– Mohr’s salt, weighing bottle, weight box, volumetric flask, funnel, distilled water, chemical balance, dilute H_2SO_4 , beakers, conical flask, funnel, burette, pipette, clamp stand, tile, $KMnO_4$ solution.

Volumetric Analysis Procedure

for Volumetric Analysis
 A typical titration starts with a beaker or flask containing a precise volume of the analyte and small amount of indicator placed underneath a calibrated burette or pipette containing the titrant.

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Comparison of the normality of the two given sodium

carbonate solutions 3.

Estimation of oxalic

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 colorimetry,
 calorimetry,
 chromatography,
 calibration procedures
 and laboratory safety.
 Through the practical
 tasks in the unit, you
 will develop proficiency
 in the quantitative
 analytical techniques
 of titration and
 colorimetry, including
 learning to calculate
 the concentration of
 solutions. You will
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